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> APPLICATION NUMBER: 60/540,566 FILING DATE: January 30, 2004

RELATED PCT APPLICATION NUMBER: PCT/US05/02745

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## PROVISIONAL APPLICATION FOR PATENT COVER SHEET

17354 B.S. PTO 60/540566 013004

This is a request for filing a PROVISIONAL APPLICATION FOR PATENT under 37 CFR 1.53 (c).

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Mailer's Name (Print/Ty	oe) Katrina	Holland	Signature	Corrini H	20 8,00	Date 12/19/03					
			INVENTOR(S)								
Given Name (first and mid	dle (if anyl)	Family N	ame or Surname	(City	and eithe	Residence r State or Foreign Country)					
Harold J.	Harold J. Schreier Baltimore, MD										
Yossi		Tai		Rchovot, Israe							
Additional invento	Additional inventors are being named on the separately numbered sheets attached hereto.										
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		ENCLOSED AP	PLICATION PARTS (ch	ck all that apply)							
x Specification Numb	er of Pages	6		CD(	s), Number						
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Application Data S	heet. See 37 C	CFR 1.76									
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PTO/SB/17 (10-03) Approved for use through 07/31/2006. OMB 0551-0032
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Effective 10/01/2003. Patent fees are subject to annual revision

X Applicant Claims small entity status. See 37 CFR 1.27

Name (Print/Type)

Marianne Fuierer

Complete if Known Application Number Filing Date January 30, 2004 Schreir, et al. First Named Inventor NA Examiner Name Art Unit Attorney Docket No. 4115-197 PRV

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### UNIVERSITY OF MARYLAND BIOTECHNOLOGY INSTITUTE (UMBI)

### DISCLOSURE AND RECORD OF INVENTION

This invention was made and/or reduced to practice in the course of work performed at the UMBI facility.

Mail completed disclosure to UMBI Office of Research & Development,

701 East Pratt Street, Suite 200, Baltimore, MD 21202

701 East Praft Street, Suite 200, Baitimore, MD 21202
Telephone (410) 385-6328, if you have questions regarding this form.

Title of Invention:

Dissimilatory Sulfate Reduction as a Process to Control Nitrate Levels in Marine Recirculated Aquaculture Systems.

Please list names of those who contributed to the inventive thought and to the final result of the invention. If a patent application is filed the Office of Research & Development will make a legal determination of inventorship.

Contributor(s):	Full Name	Title/Position	Employer	Division	Phone No.	Mail Stop
Yossi Tal		Postdoctoral Research Associate	UMBI	COMB	410-234-8875	
Harold J. Sch	reier	Associate Professor	UMBI	COMB	410-234-8874	

### Contributor(s) Permanent Home Address(es):

Name	Citizenship	Street Address	City, State, and Zip Code
Yossi Tal	Israel	4/7 Hag'ra	Rehovot, Israel
Harold J. Schreier	U.S.A.	2400 Sugarcone Road	Baltimore, MD 21209

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IMPORTANT - Check the funding source for this invention & provide the information requested:
Government Grants
Funding Agency and Grant No
Principal Investigator
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Describe
From what source do you reasonably expect future or continuing funding?
Source BARD Is this continuing funding or X future funding?
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Earliest documentation of your invention: (please provide date and identify the document)
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First Written Description:
Names of witnesses or others with knowledge of facts relating to conception:
Steve Rodgers
Reduction to Practice:
Date first model completed:
Date of operation and testing:
Place of test: Aquaculture Research Center, COMB
Results of testing:
Witnesses or others with direct knowledge of test:
Sieve Rodgers, Allen Place, Yonahan Zohar, John Stubblefield
Important: WAS ANY PROPRIETARY MATERIAL FROM OUTSIDE YOUR LABORATORY USED TO DEVELOP YOUR INVENTION (e.g., software code, cell line, antibody, DNA fragments, or chemical compound)? If yes, please describe.
ARC systems 8-1, 8-2, 8-3 and 8-4 were used to develop and test the invention.

Page 2

DESCRIPTION:
Background of the Invention Please summarize:  1) Technical problems overcome to make the invention,  2) What your invention enables people to do that couldn't be done as well before,  3) How people currently address the problem your invention addresses.
Sea Attrohed

Summary of the Invention Please include a sketch of the invention if possible (you may attach extra pages):

See Attached.

List uses of the invention - research, commercial, pilot plant, etc.; Think as broadly as possible. Identify companies that might be interested in licensing this technology.

Because the invention establishes an anaerobic environment to carry out dissimilatory sulfate reduction using organic waste produced by fish, it should be applicable for ANY marine recirculated system.

<u>Publications (including UMBI reports, mee</u> <u>DESCRIBE or <u>RELATE</u> to the invention:</u>	ting abstrac	ts, or prior patents or pa	tent application	ons) that <u>ACTU</u>	ALLY
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Page 4

Center's Director

Background and Summary:

The ability to effectively manage nutrient wastes in recirculating marine aquaculture systems is key to efficiently control these systems and limit the amount of water exchange necessary to maintain a high degree of water quality. Nitrogenous wastes are eliminated through the action of nitrifying and denitrifying biofilter units that rely on oxygen and reduced organic compounds, respectively, for their activity. In a previous invention disclosure (Provisional U.S. Patent Application No. 60/335,024) we described the stimulation of denitrification in a recirculated aquaculture bead biological filter by the addition of a slow-release carbon source. For that application, the carbon source (glucose polymers) was necessary to provide reducing equivalents to both decrease oxygen availability (making the system anaerobic) and driving the denitrification process, as exemplified in the following equation:

Glucose + 4.8 NO<sub>3</sub> + 4.8 H<sup>+</sup> 
$$\Rightarrow$$
 6 CO<sub>2</sub> + 2.4 N<sub>2</sub> (gas) + 8.4 H<sub>2</sub>O  $\Delta G^{01} = -2660 \text{ kJ}$ 

For this process, nitrate (NO<sub>2</sub>) is provided by the oxygen-dependent nitrification reaction that occurs within the aerobic compartment of the biofilter and is a result of oxidation of ammonia (NH<sub>4</sub>\*), the metabolic waste product of fish, to nitrate via a nitrite (NO<sub>2</sub>\*) intermediate, as follows:

$$NH_4 + 1.5 O_2 \rightarrow NO_2 + 2H^* + H_2O$$
  $\Delta G^0 = -274 \text{ kJ}$   
 $NO_2 + 0.5 O_2 \rightarrow NO_3$   $\Delta G^0 = -74.1 \text{ kJ}$ 

Bacteria are responsible for both nitrification and denitrification processes and we have demonstrated the presence of both nitrifying and denitrifying organisms in our system.

A major goal of developing an efficient wastewater filtration system is to decrease the frequency and amount of water exchanges necessary to maintain water quality. Decreasing this exchange not only saves on costs associated with replenishing water and salt (for marine systems) but also minimizes the level of nitrogen-rich waste that would be discharged into the environment. One significant source of water loss occurs during the removal of organic waste matter that arises from both fish excretion and uneaten food. Organic material is generally removed mechanically to avoid their consumption by bacteria, which dissipates oxygen and results in the buildup of toxic ammonia and noxious gases, such as hydrogen sulfide. This process is examplifed by the following reaction, where acetate (CH<sub>2</sub>COO) is produced from the breakdown of fatty acids and sugars and sulfate (SO<sub>4</sub>\*) is provided as a major divalent ion of sea water:

$$CH_3COO' + SO_4^{2} + 3 H^+ \rightarrow 2 CO_2 + H_2S + 2 H_2O$$
  $\Delta G^{0+} = -57.5 \text{ kJ}$ 

In this invention, we describe the potential of using the reducing capacity derived from organic waste degradation to drive denitrification in a marine recirculating fish system. We rely on coupling dissimilatory sulfate reduction to denitrification, as shown in the following equations:

Coupling dissimilatory sulfate reduction to denitrification:

$$5 \text{ H}_2\text{S} + 8 \text{ NO}_3^- \rightarrow 5 \text{ SO}_4^{-2} + \text{ N}_2 + 4 \text{H}_2\text{O}$$
  $\Delta \text{ G}^{0*} = -3721.5 \text{ kJ}$ 

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Jan, 30 2004 04:52PM P6

The system was established as follows (see accompanying diagram):

Two 4.2 m3 tanks were operated with gilthead seabream, Sparus aurata, at a density of 10-50 kg/m<sup>3</sup> and a feeding rate of 1%-1.6% body weight/day. The tanks were connected to a 2 m3 nitrifying moving bed bioreactor (MBB) with 1m3 of polyethylene beads with specific surface area of 500 m<sup>2</sup>/m<sup>3</sup> (4.86 cm<sup>2</sup>/bead). A flow rate of 8 m<sup>3</sup>/hr was set to enable two exchanges of tank water per hour through the filter. Attached to this tank as a side loop was a 0.3 m<sup>3</sup> cylindrical up-flow fixed bed biofilter filled with 0.2 m<sup>3</sup> of polyethylene beads for denitrification. This anaerobic biofilter component was set with a low flow rate of 0.1 m3/hr (this set-up was described in Provisional U.S. Patent Application No. 60/335,024). Sludge (organic waste material) collection was carried out through a drum screen filter with backwash system that used tank water. Sludge and backwash water were collected in a 0.3 m3 rectangular tank with 0.1 m3 beads that provided a means for solids removal as well as substrate for bacterial colonization. Water from the sludge tank was pumped back into the system via the anaerobic biofilter and high-density sludge was collected and removed weekly. Ammonia, nitrite and nitrate levels as well as sulfate, sulfide, COD, pH and oxygen were monitored in each compartment. Our results indicated that the sludge/denitrifying filter sequence was very effective in stimulating nitrate removal. The denitrifying compartment removed as much as 40-70% of the nitrate load introduced from the system, which allowed for overall system nitrate concentrations to be maintained between 35-65 mg (NO<sub>3</sub>-N)/l during operation and enabled daily water exchange to be less than 1% of total system volume. Stimulation of nitrate reducing activity was likely due, in part, to the presence of dissimilatory sulfate reduction activity that occurred during sludge waste decomposition; sulfide levels within the sludge compartment were as high as 60-80 mg/l and the redox potential reached values lower than -500 mV. On the other hand, effluent water from the denirifying compartment showed no measurable sulfide and a redox potential between -50 and +100 mV.

We believe that our invention allows for denitrification to be driven by the redox gradient between system compartments resulting in sulfate reduction (sulfide production), which, in turn, was utilized for nitrate reduction. This process has wide application for marine recirculating systems, where sulfate concentration is not a limiting factor and minimizing water exchange is critical.

Results for first 49 days of operation are shown in Figure 1.

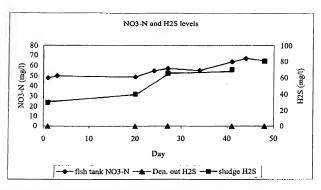


Figure 1. Nitrate-N levels for the tank system and levels for sulfide measured at the denitrifying compartment outlet (Den. Out) and sludge compartment.

# Sulfide/Denitrification Coupled System

